



Mark Scheme (Results)

Summer 2021

Pearson Edexcel International Advanced Level
In Chemistry (WCH14)

Paper 01: Rates, Equilibria and Further Organic
Chemistry

Section A (multiple choice)

Question Number	Correct Answer	Mark
1	The only correct answer is D (SO₂) <i>A is incorrect as although it has four atoms, it has ten electrons</i> <i>B is incorrect as it has two atoms and two electrons</i> <i>C is incorrect as it has two atoms and only fourteen electrons</i>	1

Question Number	Correct Answer	Mark
2	The only correct answer is A (– 198.8) <i>B is incorrect as number of moles of NH₃ and H₂ have not been considered</i> <i>C is incorrect as number of moles of NH₃ and H₂ have not been considered and the expression to find the standard entropy of the system is the wrong way round</i> <i>D is incorrect as expression to find the standard entropy of the system is the wrong way round</i>	1

Question Number	Correct Answer	Mark
3	The only correct answer is C (enthalpy change of formation of Na₂SO₄) <i>A is incorrect as lattice energy is used to find the enthalpy change of solution</i> <i>B is incorrect as enthalpy change of hydration is used to find the enthalpy change of solution</i> <i>D is incorrect as enthalpy change of hydration is used to find the enthalpy change of solution</i>	1

Question Number	Correct Answer	Mark
4 (a)	The only correct answer is C ($0.1 \text{ mol dm}^{-3} \text{ HCl}$) <i>A is incorrect as final pH would be greater than 2 (weak acid)</i> <i>B is incorrect as final pH would be greater than 2 (weak acid)</i> <i>D is incorrect as final pH would be less than 1 (strong acid)</i>	1

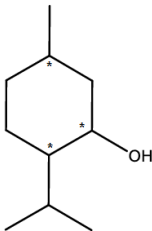
Question Number	Correct Answer	Mark
4 (b)	The only correct answer is A (NH_3) <i>B is incorrect as strong base so vertical section would begin at a higher pH / curve has a buffer region</i> <i>C is incorrect as strong base so vertical section would begin at a higher pH / curve has a buffer region</i> <i>D is incorrect as strong base so vertical section would begin at a higher pH / curve has a buffer region</i>	1

Question Number	Correct Answer	Mark
4 (c)	The only correct answer is C (3) <i>A is incorrect as only methyl orange, bromophenol blue and bromocresol green would change colour in the vertical section of the 'curve'</i> <i>B is incorrect as only methyl orange, bromophenol blue and bromocresol green would change colour in the vertical section of the 'curve'</i> <i>D is incorrect as methyl orange, bromophenol blue and bromocresol green would change colour in the vertical section of the 'curve'</i>	1

Question Number	Correct Answer	Mark
5	The only correct answer is D (S_N1 ; Two steps in mechanism) <i>A is incorrect as the halogenoalkane is tertiary so mechanism would be S_N1 which has two steps</i> <i>B is incorrect as although the mechanism has two steps the halogenoalkane is tertiary so mechanism would be S_N1</i> <i>C is incorrect as although the mechanism is S_N1, it would have two steps</i>	1

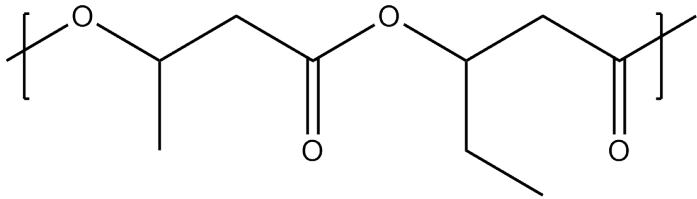
Question Number	Correct Answer	Mark
6	The only correct answer is D (Step 2 is the rate determining step, the overall order is 3) <i>A is incorrect as Step 3 is fast</i> <i>B is incorrect as Step 3 is fast</i> <i>C is incorrect as the overall order is 3</i>	1

Question Number	Correct Answer	Mark
7	The only correct answer is C ($-\text{gradient} \times R$) <i>A is incorrect the Arrhenius equation has been rearranged incorrectly</i> <i>B is incorrect as the gradient of the graph is negative, so this expression would give a negative value for an activation energy</i> <i>D is incorrect as the gradient of the graph is negative, so this expression would give a negative value for an activation energy</i>	1

Question Number	Correct Answer	Mark
8 (a)	The only correct answer is C (3)  <i>A is incorrect as menthol has 3 chiral carbon atoms</i> <i>B is incorrect as menthol has 3 chiral carbon atoms</i> <i>D is incorrect as menthol has 3 chiral carbon atoms</i>	1

Question Number	Correct Answer	Mark
8 (b)	The only correct answer is B (Q) <i>A is incorrect as this carbon would produce a peak between 0 and 60 ppm</i> <i>C is incorrect as this carbon would produce a peak between 0 and 60 ppm</i> <i>D is incorrect as this carbon would produce a peak between 0 and 60 ppm</i>	1

Question Number	Correct Answer	Mark
8 (c)	The only correct answer is B (Two) <i>A is incorrect as the oxidation product is a ketone, so would not react with PCl_5</i> <i>C is incorrect as the oxidation product is a ketone, so would not react with Fehling's solution</i> <i>D is incorrect as the oxidation product is a ketone, so would not react with PCl_5 but would react with 2,4-dinitrophenylhydrazine</i>	1

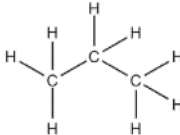
Question Number	Correct Answer	Mark
9 (a)	The only correct answer is B  <i>A is incorrect as the repeat unit has an extra oxygen</i> <i>C is incorrect as there is an extra carbon at the left-hand end of the repeat unit</i> <i>D is incorrect as the repeat unit has an extra oxygen and the structure is incorrect</i>	1

Question Number	Correct Answer	Mark
9 (b)	The only correct answer is B (hydrolysis) <i>A is incorrect as condensation is the reaction when the polymer forms</i> <i>C is incorrect as hydration is the addition of water to a C=C bond</i> <i>D is incorrect as hydrogen has not been added in a reduction reaction</i>	1

Question Number	Correct Answer	Mark
10	The only correct answer is D (CH_3COCl) <i>A is incorrect as the reaction with ketone would NOT form an N-substituted amide</i> <i>B is incorrect as any reaction with the carboxylic acid would be too slow at RT</i> <i>C is incorrect as any reaction with the ester would be too slow at RT</i>	1

Question Number	Correct Answer	Mark
11(a)	The only correct answer is B (68 mm) <i>A is incorrect as it is a factor of 10 too large</i> <i>C is incorrect as it is the distance moved by the amino acids</i> <i>D is incorrect as it is the expression for R_f has been inverted</i>	1

Question Number	Correct Answer	Mark
11 (b)	The only correct answer is A (argon) <i>B is incorrect as the carrier gas must be inert</i> <i>C is incorrect as the carrier gas must be inert</i> <i>D is incorrect as the carrier gas must be inert</i>	1

Question Number	Correct Answer	Mark
12	The only correct answer is C  <i>A is incorrect as the molar mass to 4 dp is 44.0265</i> <i>B is incorrect as the molar mass to 4 dp is 44.0265</i> <i>D is incorrect as the molar mass to 4 dp is 43.9898</i>	1

Question Number	Correct Answer	Mark
13	The only correct answer is D (8) <i>A is incorrect as the number of optical isomers = 2^n, where n = number of chiral centres</i> <i>B is incorrect as the number of optical isomers = 2^n, where n = number of chiral centres</i> <i>C is incorrect as the number of optical isomers = 2^n, where n = number of chiral centres</i>	1

Question Number	Correct Answer	Mark
14	The only correct answer is D (Structure D) <i>A is incorrect as it is identical to B and C</i> <i>B is incorrect as it is identical to A and C</i> <i>C is incorrect as it is identical to A and B</i>	1

(Total for Section A = 20 marks)

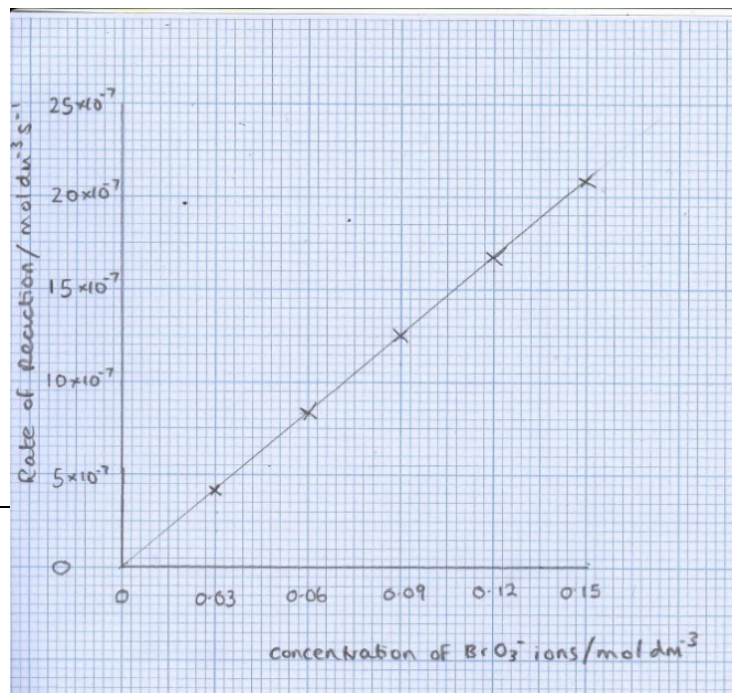
Section B

Question Number	Acceptable Answers	Additional Guidance	Mark
15 (a)	- correct expression for $\Delta S_{\text{surroundings}}$ (1) - correct evaluation and correct units and correct sign (1)	<u>Example of calculation</u> $-\Delta H \div T = -25.7 \div 298$ $-0.086242 \text{ kJ K}^{-1} \text{ mol}^{-1} /$ $-86.242 \text{ J K}^{-1} \text{ mol}^{-1}$ Ignore SF except 1 SF Correct answer with no working scores (2) Allow TE in M2 for use of $\Delta H \div T$ Comment Mark value first – if correct, with units and sign award 2 marks For units allow $\text{kJ K}^{-1} \text{ mol}^{-1} / \text{J K}^{-1} \text{ mol}^{-1}$	2

Question Number	Acceptable Answers	Additional Guidance	Mark
15(b)	An explanation that makes reference to: ΔS_{system} must be positive (1) $\Delta S_{\text{system}} > 86.24 \text{ J mol}^{-1}$ / answer to (a) (1) (as compound does dissolve) ΔS_{total} is > 0 / positive (1)	Allow 'ΔS_{system} is more positive' Allow $T\Delta S_{\text{system}}$ is greater in magnitude / more negative than ΔH ΔG is negative If answer to (a) is positive , then M1 and M2 will be ΔS_{system} could be positive or negative ΔS_{system} smaller in magnitude than answer to (a) / $T\Delta S_{\text{system}}$ is greater than ΔH	3

(Total for Question 15 = 5 marks)

Question Number	Acceptable Answers	Additional Guidance	Mark
16(a)(i)	- rate against concentration graph with axes labelled, inc. units - suitable scale chosen including the origin - all points plotted correctly and straight line of best fit.	(1) Do not award M1 if axes are the other way around Allow $[\text{BrO}_3^-] / \text{mol dm}^{-3}$ Ignore 'initial' (1) Points cover at least half available space in both directions (1) Allow $\pm \frac{1}{2}$ a square Allow if line does not extend to the origin Do not award M3 if scale is non-linear	3



Question Number	Acceptable Answers	Additional Guidance	Mark
16(a)(ii)	justification of first order	(First order with respect to BrO_3^-) as straight line (through origin / 0,0) Allow line with constant gradient Allow rate is (directly) proportional to concentration Allow use of data from graph to justify order Do not award 'constant half life'	1

Question Number	Acceptable Answers	Additional Guidance	Mark
16(b)(i)	- deduce order wrt Br^- ions (1) - deduce order wrt H^+ ions (1)	1st order 2nd order	2

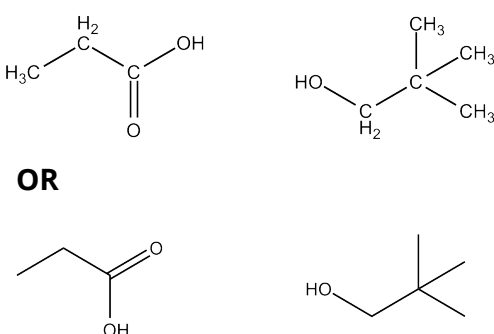
Question Number	Acceptable Answers	Additional Guidance	Mark
16(b)(ii)	rate equation shown	$\text{rate} = k[\text{BrO}_3^-][\text{Br}^-][\text{H}^+]^2$ Allow TE from (b)(i)	1

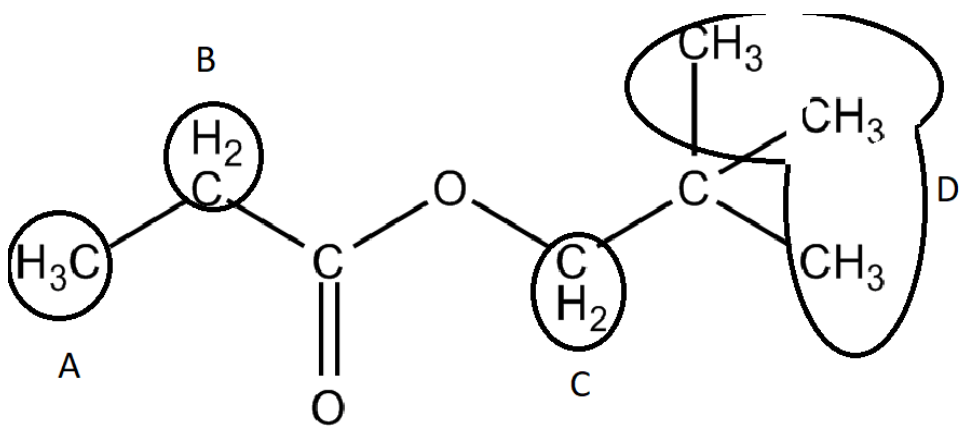
Question Number	Acceptable Answers	Additional Guidance	Mark
16(b)(iii)	- rearrangement of rate equation - evaluation of k - units for k	(1) <u>Example of calculation</u> $k = \text{rate}/[\text{BrO}_3^-][\text{Br}^-][\text{H}^+]^2 /$ $k = 1.52 \times 10^{-5} / (0.062 \times 0.21 \times 0.4^2)$ (1) 7.2965×10^{-3} ignore SF except 1SF M1 can be subsumed within award of M2 (1) $\text{dm}^9 \text{mol}^{-3} \text{s}^{-1}$ allow in any order Correct answer with no working scores (3) TE on (b)(ii) Allow use of data from Run 2 or Run 3	3

Question Number	Acceptable Answers	Additional Guidance	Mark
16(c)	An answer that makes reference to: - reactants adsorb onto palladium/catalyst surface - this weakens bonds in reactants - products then desorb (from catalyst surface)	Allow bromate ((V)) ions for reactants Allow 'bond/bind onto catalyst surface' Do not award absorb Ignore comments related to orientation Allow 'products de-adsorb' / products released (from catalyst surface) If no other mark is awarded allow one for: reaction follows an alternative pathway / route / mechanism of lower activation energy	3

(Total for Question 16 = 13 marks)

Question Number	Acceptable Answers	Additional Guidance			Mark	
17(a)	calculation of moles of C, H and O (1)		element	moles	ratio	2
		C	66.7÷12 =5.56	5.56÷1.3875 = 4		
		H	11.1÷1 = 11.1	11.1÷1.3875 = 8		
		O	22.2÷16 =1.3875	1.3875÷1.3875 = 1		
	calculation of ratio and identify that ratio matches molecular formula (1)	Ratio C ₄ H ₈ O matches C ₈ H ₁₆ O ₂				
	OR					
	calculate molar mass of Y (1)	Molar mass = 144 (g mol ⁻¹)				
	calculate % of each element (1)	C=96÷144x100 = 66.7% H=16÷144x100 = 11.1% O=32÷144x100 = 22.2%				

Question Number	Acceptable Answers	Additional Guidance	Mark
17(b)(i)	2,2-dimethylpropyl propanoate (2)	Any name with '-propyl propanoate' scores 1 propyl-2,2-dimethyl propanoate scores 1 2,2-dimethyl propanoate scores 1 2,2-dimethylpropyl ethanoate scores 1	2
Question Number	Acceptable Answers	Additional Guidance	Mark
17(b)(ii)	 OR	Both structures required for mark Allow structures of propanoyl chloride / propanoic anhydride Allow any combination of correct skeletal, structural or displayed formulae. Ignore names even if incorrect Do not award connectivity to hydroxyl group via H atom	1

Question Number	Acceptable Answers	Additional Guidance	Mark
17(c)(i)		Labels B C and D can be used interchangeably as long as the three proton environments are identified correctly. Allow 3 methyl groups to be circled individually but with a single label / labels pointing to all 3	1

Question Number	Acceptable Answers	Additional Guidance	Mark															
17(c)(ii)	<table><tr><th>Hydrogen environment</th><th>Splitting pattern of peak</th><th>Relative peak area</th></tr><tr><td>(A)</td><td>(triplet)</td><td>(3)</td></tr><tr><td>B</td><td>quartet</td><td>2</td></tr><tr><td>C</td><td>singlet</td><td>2</td></tr><tr><td>D</td><td>singlet</td><td>9</td></tr></table>	Hydrogen environment	Splitting pattern of peak	Relative peak area	(A)	(triplet)	(3)	B	quartet	2	C	singlet	2	D	singlet	9	1 mark for each row. But If two or more rows are incorrect then award whichever of these alternatives is higher Allow 2 marks for 3 correct splitting patterns. OR Allow 1 mark for 3 correct peak areas. OR Allow 1 mark for correct row marked consequentially on the labelling in 17(c)(i)	3
	Hydrogen environment	Splitting pattern of peak	Relative peak area															
	(A)	(triplet)	(3)															
	B	quartet	2															
	C	singlet	2															
	D	singlet	9															
	Note : allow 'quadruplet' as alternative for quartet / 																	

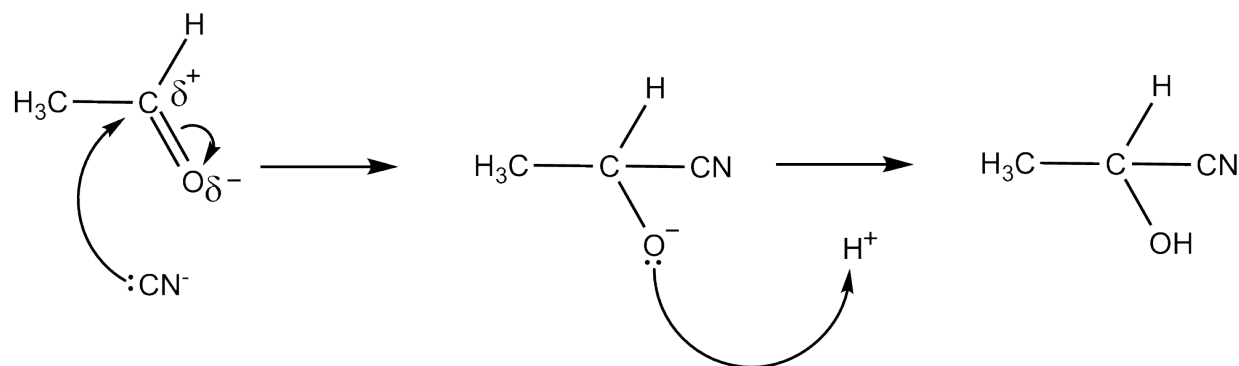
(Total for Question 17 = 9 marks)

Question Number	Acceptable Answers	Additional Guidance	Mark																				
18	This question assesses the student’s ability to show a coherent and logically structured answer with linkages and fully sustained reasoning. Marks are awarded for indicative content and for how the answer is structured and shows lines of reasoning. The following table shows how the marks should be awarded for indicative content. <table><tr><th>Number of indicative marking points seen in answer</th><th>Number of marks awarded for indicative marking points</th></tr><tr><td>6</td><td>4</td></tr><tr><td>5-4</td><td>3</td></tr><tr><td>3-2</td><td>2</td></tr><tr><td>1</td><td>1</td></tr><tr><td>0</td><td>0</td></tr></table> The following table shows how the marks should be awarded for structure and lines of reasoning <table><tr><th></th><th>Number of marks awarded for structure of answer and sustained lines of reasoning</th></tr><tr><td>Answer shows a coherent logical structure with linkages and fully sustained lines of reasoning demonstrated throughout</td><td>2</td></tr><tr><td>Answer is partially structured with some linkages and lines of reasoning</td><td>1</td></tr><tr><td>Answer has no linkages between points and is unstructured</td><td>0</td></tr></table> Indicative Points	Number of indicative marking points seen in answer	Number of marks awarded for indicative marking points	6	4	5-4	3	3-2	2	1	1	0	0		Number of marks awarded for structure of answer and sustained lines of reasoning	Answer shows a coherent logical structure with linkages and fully sustained lines of reasoning demonstrated throughout	2	Answer is partially structured with some linkages and lines of reasoning	1	Answer has no linkages between points and is unstructured	0	Guidance on how the mark scheme should be applied: The mark for indicative content should be added to the mark for lines of reasoning. For example, a response with four indicative marking points that is partially structured with some linkages and lines of reasoning scores 4 marks (3 marks for indicative content and 1 mark for partial structure and some linkages and lines of reasoning). If there were no linkages between the points, then the same indicative marking points would yield an overall score of 3 marks (3 marks for indicative content and zero marks for linkages).	6
Number of indicative marking points seen in answer	Number of marks awarded for indicative marking points																						
6	4																						
5-4	3																						
3-2	2																						
1	1																						
0	0																						
	Number of marks awarded for structure of answer and sustained lines of reasoning																						
Answer shows a coherent logical structure with linkages and fully sustained lines of reasoning demonstrated throughout	2																						
Answer is partially structured with some linkages and lines of reasoning	1																						
Answer has no linkages between points and is unstructured	0																						

	• IP1 Bonding in sodium chloride is (almost) 100% ionic bonds and as the theoretical and Born-Haber values are (very) similar • IP2 Bonding in magnesium iodide has some covalent character and as theoretical and Born-Haber values are (significantly) different • IP3 Anion is (more) polarised in magnesium iodide (than sodium chloride) • IP4 Magnesium ion has a greater charge density (than sodium ion), so greater polarising power • IP5 Iodide ion is larger (than chloride ion), so is more easily polarised • IP6 Magnesium iodide has stronger bonding than sodium chloride because the charge on the magnesium ion is twice as large (as the charge on the sodium ion)	If neither IP1 or IP2 scored can get 1IP for Bonding in sodium chloride is (almost) 100% ionic bonds and bonding in magnesium iodide has some covalent character ALLOW Magnesium ion has a greater charge/smaller than sodium ion, so greater polarising power polarisation must be mentioned at least once in IP3, IP4 and IP5 Penalise use of 'atoms' instead of ions once only in IP3 IP4 and IP5 Penalise lack of comparative language once only in IP4, IP5 and IP6 Allow magnesium iodide has stronger bonding (than expected) due to polarisation / covalent character Allow both compounds have strong bonds as large amounts of energy needed to break up lattice / released when lattice forms / needed to break many strong bonds	
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(Total for Question 18 = 6 marks)

Question Number	Acceptable Answers	Additional Guidance	Mark
19(a)(i)	- arrow from lone pair on carbon of cyanide ion to carbonyl carbon (1) - dipoles on carbon and oxygen in carbonyl bond and arrow from carbonyl bond to oxygen or just beyond (1) - structure of intermediate, including charge (1) - arrow from lone pair of oxygen in intermediate to hydrogen ion / H in HCN (1)	Penalise absence of lone pair only once in M1, M3 and M4 If HCN used to protonate in step 2, dipole on HCN and curly arrow to break HCN bond are not required Ignore product	4



Question Number	Acceptable Answers	Additional Guidance	Mark
19(a)(ii)	The prediction is incorrect because - ethanal is planar around the carbonyl carbon atom / planar around the CHO (1) (so in Step 1) the (carbonyl) carbon can be attacked from above or below (1) - hence both stereoisomers (of intermediate / product) will form in equal amounts or so product mixture is racemic / rotates the plane of plane-polarised light equally in both directions (1)	Accept planar at the site of the nucleophilic attack / planar about C=O Do not award planar molecule / cation / intermediate Allow attack from any direction / either side Ignore 'has no effect on the plane of plane-polarised light' Ignore comments related to SN1 or SN2 If no other mark scored allow 1 mark for idea that product will rotate plane of plane polarised light as it has a chiral centre / carbon	3

Question Number	Acceptable Answers	Additional Guidance	Mark
19(a)(iii)	- hydrolysis (1) (dilute) hydrochloric acid / HCl(aq) (1) - heat (under reflux) / reflux (1) $\text{CH}_3\text{CH}(\text{OH})\text{CN} + 2\text{H}_2\text{O} + \text{H}^+ \rightarrow \text{CH}_3\text{CH}(\text{OH})\text{COOH} + \text{NH}_4^+$ (1) OR $\text{CH}_3\text{CH}(\text{OH})\text{CN} + 2\text{H}_2\text{O} \rightarrow \text{CH}_3\text{CH}(\text{OH})\text{COOH} + \text{NH}_3$ OR $\text{CH}_3\text{CH}(\text{OH})\text{CN} + 2\text{H}_2\text{O} + \text{HCl} \rightarrow \text{CH}_3\text{CH}(\text{OH})\text{COOH} + \text{NH}_4\text{Cl}$ OR $\text{CH}_3\text{CH}(\text{OH})\text{CN} + \text{H}_2\text{O} + \text{OH}^- \rightarrow \text{CH}_3\text{CH}(\text{OH})\text{COO}^- + \text{NH}_3$ and $\text{CH}_3\text{CH}(\text{OH})\text{COO}^- + \text{H}^+ \rightarrow \text{CH}_3\text{CH}(\text{OH})\text{COOH}$	Allow any strong acid by name or formula Allow sodium hydroxide followed by any (strong) acid Ignore conc / concentrated Allow 'boil' for heat Ignore 'warm' Allow NaOH for OH^- Allow HCl for H^+ Ignore state symbols even if incorrect	4

Question Number	Acceptable Answers	Additional Guidance	Mark
19(b)	• $\text{CH}_3\text{CH}(\text{OH})\text{COOH} + \text{NaHCO}_3 \rightarrow \text{CH}_3\text{CH}(\text{OH})\text{COO}^-\text{Na}^+ + \text{H}_2\text{O} + \text{CO}_2$ OR $\text{H}^+ + \text{HCO}_3^- \rightarrow \text{H}_2\text{O} + \text{CO}_2$	Allow $\text{CH}_3\text{CH}(\text{OH})\text{COONa}$ Allow H_2CO_3 Ignore state symbols even if incorrect Do not award if covalent bond shown between O and Na	1

Question Number	Acceptable Answers	Additional Guidance	Mark
19(c)(i)	(large concentration of) HCO_3^- react with (extra) H^+ ions (1) - equilibrium 1 moves to the RHS to keep concentration of H^+ ions constant / H_2CO_3 forms to keep concentration of H^+ ions constant (1) - equilibrium 2 moves to RHS to form CO_2 (which can be excreted from the body) / H_2CO_3 then forms CO_2 (and water) (1)	Allow ratio of $[\text{HCO}_3^-]$ to $[\text{H}_2\text{CO}_3]$ remains constant / ratio of [salt] to [acid] remains constant Allow H_3O^+ for H^+ Allow equilibrium 1 moves to the RHS to remove excess H^+ ions / H_2CO_3 forms to remove excess H^+ ions If no reference to H^+ and CO_2 in M2 and M3 but direction of movement of equilibria are correct in both cases, allow 1 mark	3

Question Number	Acceptable Answers	Additional Guidance	Mark
19(c)(ii)	- calculation of $[H^+] / [H_3O^+]$ (1) K_a expression (1) - rearrangement of K_a expression and calculation of $[HCO_3^-] : [H_2CO_3]$ (1) OR - calculation of pK_a (1) - Henderson Hasselbach expression (1) - rearrangement of K_a expression and calculation of $[HCO_3^-] : [H_2CO_3]$ (1)	$[H^+] = 10^{-7.41} / = 3.8905 \times 10^{-8}$ $K_a = \frac{[HCO_3^-][H^+]}{[H_2CO_3]}$ Allow $[H_3O^+]$ in K_a Do not award $[H_2O]$ in K_a expression $[HCO_3^-] : [H_2CO_3]$ $= 4.5 \times 10^{-7} \div 3.8905 \times 10^{-8} = 11.567 : 1 = 11.6 (: 1)$ Ignore SF except 1 Allow correct rounding of $[H^+]$ to 3.9×10^{-8} Allow $1 : 0.086444$ if it's clear that 1 relates to $[HCO_3^-]$ $pK_a = -\log 4.5 \times 10^{-7} = 6.3468$ $pH = pK_a + \log([HCO_3^-] \div [H_2CO_3])$ $7.41 - 6.3468 = \log([HCO_3^-] \div [H_2CO_3])$ $[HCO_3^-] : [H_2CO_3] = 11.567 (: 1)$ Correct answer with no working scores (3) If final answer close, check for and allow correct rounding	3

(Total for Question 19 = 18 marks)

(Total for Section B = 51 marks)

Question Number	Acceptable Answers	Additional Guidance	Mark
20(a)(i)	When the pressure is increased - equilibrium moves to RHS and yield (of chlorine) increases (1) - as fewer gas molecules on the RHS (5:4) (1) K_p remains constant (1)	Marking points are independent Allow 'forward reaction favoured so yield (of chlorine) increases' If numbers are given they must be correct Allow use of 4/5 ratio to justify decrease in quotient / greater increase in denominator as total pressure increases, (so eqm moves (to RHS) to restore K_p) Allow 'change in pressure has no effect on value for K_p'	3

Question Number	Acceptable Answers	Additional Guidance	Mark
20(a)(ii)	When the temperature increases - equilibrium moves to LHS as (forward) reaction is exothermic (1) K_p decreases and so yield (of chlorine) decreases (1)	Marking points are independent Allow reaction moves in endothermic direction Allow increase in T reduces ΔS_{surr} and hence ΔS_{total} decreases	2

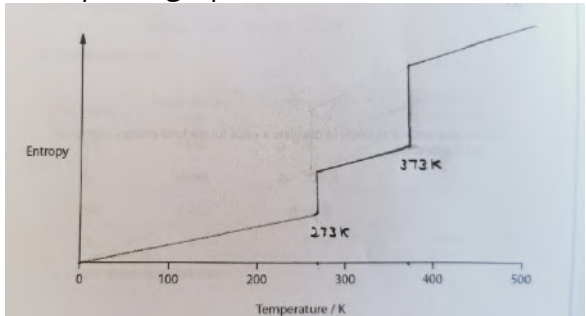
Question Number	Acceptable Answers	Additional Guidance	Mark
20(a)(iii)	When a catalyst is used - rate of backward and forward reactions increases by same amount (1) - so K_p and yield (of chlorine) is unchanged (1)		2

Question Number	Acceptable Answers	Additional Guidance	Mark																								
20(b)(i)	<table border="1"> <thead> <tr> <th>Substance</th><th>Initial amount / mol</th><th>Equilibrium amount / mol</th><th>Mole fraction at equilibrium</th></tr> </thead> <tbody> <tr> <td>HCl</td><td>0.850</td><td>0.350</td><td>0.26415</td></tr> <tr> <td>O₂</td><td>0.600</td><td>0.475</td><td>0.35849</td></tr> <tr> <td>H₂O</td><td>0</td><td>0.250</td><td>0.18868</td></tr> <tr> <td>Cl₂</td><td>0</td><td>0.250</td><td>0.189</td></tr> <tr> <td colspan="3">Total moles at equilibrium = 1.325</td><td></td></tr> </tbody> </table> All values correct scores (3) M1 1 correct equilibrium amount (1) M2 other 2 correct equilibrium amounts (1) M3 Consequential total moles and mol fraction (1)	Substance	Initial amount / mol	Equilibrium amount / mol	Mole fraction at equilibrium	HCl	0.850	0.350	0.26415	O ₂	0.600	0.475	0.35849	H ₂ O	0	0.250	0.18868	Cl ₂	0	0.250	0.189	Total moles at equilibrium = 1.325				For mole fractions allow e.g. $0.350 \div 1.325$ allow correct rounding Ignore SF except 1 SF	3
Substance	Initial amount / mol	Equilibrium amount / mol	Mole fraction at equilibrium																								
HCl	0.850	0.350	0.26415																								
O ₂	0.600	0.475	0.35849																								
H ₂ O	0	0.250	0.18868																								
Cl ₂	0	0.250	0.189																								
Total moles at equilibrium = 1.325																											

Question Number	Acceptable Answers	Additional Guidance	Mark
20(b)(ii)	$K_p = p(\text{H}_2\text{O})^2 p(\text{Cl}_2)^2 \div p(\text{HCl})^4 p(\text{O}_2)$	Ignore parentheses Do not award square brackets	1

Question Number	Acceptable Answers	Additional Guidance	Mark										
20(b)(iii)	- mole fractions converted to partial pressure (1) - final value for K_p given to 2 or 3SF (1) - correct units given (1)	Example of calculation allow TE from 20b(i) <table><tr><th>substance</th><th>pp</th></tr><tr><td>HCl</td><td>0.39623</td></tr><tr><td>O₂</td><td>0.53774</td></tr><tr><td>H₂O</td><td>0.28302</td></tr><tr><td>Cl₂</td><td>0.28302</td></tr></table> Allow e.g. for pp(HCl) ; 0.26415 x 1.5 $\frac{(0.28302)^2(0.28302)^2}{(0.39623)^4(0.53770)}$ = 0.48407 (Note = 0.48408 if no rounding) = 0.48 / 0.484 No TE for M2 for incorrect expression Check final answer if close, and allow if correct rounding used in working atm⁻¹ allow TE for M3 from incorrect expression in (b)(ii)	substance	pp	HCl	0.39623	O ₂	0.53774	H ₂ O	0.28302	Cl ₂	0.28302	3
substance	pp												
HCl	0.39623												
O ₂	0.53774												
H ₂ O	0.28302												
Cl ₂	0.28302												

Question Number	Acceptable Answers	Additional Guidance	Mark
20(b)(iv)	- recall of expression for ΔS_{total} (1) - calculation of ΔS_{total} (1)	$\Delta S_{\text{total}} = R \ln K$ $= 8.31 \times -0.726$ $= -6.033 \text{ (J K}^{-1} \text{ mol}^{-1}\text{)}$ Allow TE / rounded value from (iii) No TE for M2 from incorrect expression Ignore SF except 1 SF Ignore units even if incorrect NOTE $\Delta S_{\text{total}} = -6.0289$ if no rounding from (b)(iii) $\Delta S_{\text{total}} = -6.0993$ if 0.48 used from b(iii)	2

Question Number	Acceptable Answers	Additional guidance	Mark
20(c)	- general shape of increase from left to right ALLOW straight line (1) - two vertical stages for melting and boiling (1) - include the use of 273K for melting and 373K for boiling temperature either by labelling or position on x axis (1)	Allow horizontal sections allowed between phase changes for M1 M3 is independent of M2, providing a line is drawn <u>Example of graph</u> 	3

(Total for Question 20 = 19 marks)

(Total for Section C = 19 marks)

Total for Paper = 90 marks